Summer Reading for AP/Honors:
Social Studies, Science, and English

Social Studies

9th grade World Civilizations and Culture I/Honors: Click here for assignment
This Fleeting World: A Short History of Humanity by David Christian (Great Barrington, MA: Berkshire

All students should copy the reading guide and assigned questions here.

Honors World Civilizations and Cultures II: Click here for assignment
Mathabane, Mark. Kaffir Boy: The True story of a Black Youth’s Coming of Age in Apartheid South Africa.

AP Modern European History: Click here for assignment

AP World History:
Please see Mr. Lofaro after June 1 for your assignment.

AP United States History:

Larson, Eric. The Devil in the White City: Murder, Magic, and Madness at the Fair That Changed

AP Economics: Click here for assignment
Buchholz, Todd G. New Ideas From Dead Economists: An Introduction to Modern Economic Thought.


AP Psychology

Phineas Gage: A Gruesome but True Story about Brain Science by John Fleischman

AP Government:

REQUIRED READINGS:

**RECOMMENDED READINGS FOR AP GOVERNMENT:**


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**Science**

**AP Biology:** [Click here for assignment](#)

**AP Environmental Science:** [Click here for assignment](#)
Summer Reading for English

9th Honors
Autobiography of My Dead Brother by Walter Dean Myers
Haroun and the Sea of Stories by Salman Rushdie  Chapter One

10th Honors World Lit I
The Heights of Macchu Picchu- Pablo Neruda
Snow Flower and the Secret Fan- Lisa See

11th grade Honors English
The Crucible by Arthur Miller

11th AP Language and Composition-
“Where Are You Going? Where Have You Been?” by Joyce Carol Oates
“Letter from Birmingham Jail” by Martin Luther King
Of Mice and Men by John Steinbeck
To Kill a Mockingbird by Harper Lee

12 AP Literature and Composition  For each reading, fill out the data sheet found here.
Jane Eyre by Charlotte Bronte
The Namesake by Jhumpa Lahiri
Hedda Gabler by Henrik Ibsen

12th grade Honors English
The Namesake by Jhumpa Lahiri
Dear prospective AP Chemistry student,

Welcome to AP Chemistry! The entire course focuses on the six “Big Ideas” below:

Big Idea 1: The chemical elements are fundamental building materials of matter, and all matter can be understood in terms of arrangements of atoms. These atoms retain their identity in chemical reactions.

Big Idea 2: Chemical and physical properties of materials can be explained by the structure and the arrangement of atoms, ions, or molecules and the forces between them.

Big Idea 3: Changes in matter involve the rearrangement and/or reorganization of atoms and/or the transfer of electrons.

Big Idea 4: Rates of chemical reactions are determined by details of the molecular collisions.

Big Idea 5: The laws of thermodynamics describe the essential role of energy and explain and predict the direction of changes in matter.

Big Idea 6: Any bond or intermolecular attraction that can be formed can be broken. These two processes are in a dynamic competition, sensitive to initial conditions and external perturbations.

You may (and should) view the entire College Board AP Chemistry syllabus at this link.

I encourage you to check out Appendix B found in the link above. This appendix contains all of the equations and constants that you will need to use during the course and in the AP exam. Anything on this formula sheet is fair game for the exam.

Our AP Chemistry course has very high expectations and there is no time to waste. To ensure that you are successful in the class and on the AP Exam next May, we need to start before class begins. I hope that these assignments will be helpful in easing the transition back to school in September.

Required Course Materials:

1. A Large 3 ring binder - solely for AP Chemistry.

2. A four function calculator - nothing fancy! You can find one at the dollar store.

3. Pencils, Pens and Colored Pencils
AP Chemistry Summer Assignment

Students can access the textbook at this web address:


Assignment:

1. Carefully read and outline in your notes Chapters 1, 2, and chapter 3 sections 3.1 to 3.8 inclusive.

Chapter 1:

Know the common base units in the SI system

Understand uncertainty in measurement

DEFINITELY know the rules of significant figures

Be able to determine density and convert between the 3 temperature systems (°F, °C and K)

Know the classification system for matter

Be able to describe the techniques of chromatography, distillation, and filtration

Be able to characterize changes as physical or chemical.

Read about dimensional analysis and do the chapter end problems, but I will teach you my own way of doing it

Answer end of chapter questions (to be handed in day 1 of class): 2,3, 4,8,9,11,14,

You can look at ALL of the blue-numbered problems from 17 to 93, but you must complete and hand in black-numbered problems: 22, 24, 26, 28, 30, 32, 36, 38, 42, 44, 46, 50, 52, 54, 68, 70, 72, 74, 76, 80, 90, 94

Chapter 2:

Know and understand the laws of conservation of mass, definite composition, and multiple proportion.

Be able to state the four parts of Dalton’s Atomic Theory

Be able to draw representations of, and explain the experiments of JJ Thomson and Ernest Rutherford that led to the discovery of the electron, proton, and nucleus of the atom.
Know the locations of: metals, metalloids and non-metals, alkali metals, alkaline earth metals, halogens and noble gases on the periodic table.

Know the formulas, charges and names of the ions in Tables 2.3, 2.4, and 2.5

Be able to name compounds given a formula, or derive formulas given the names.

Answer end of chapter questions (to be handed in day 1 of class) 2, 4, 5, 9, 11, 20, 24, 32, 44, 46, 50, 52, 64, 66, 68, 70, 94

Chapter 3:

Understand how the mass spectrometer works.

Know how the amu is derived

Know how to determine the average weighted atomic mass of an element.

Know what the mole represents

Be able to determine molar mass and percent composition from a formula

Given percentage data, be able to derive an empirical formula.

Be able to balance chemical equations

Answer end of chapter questions: 1, 20, 22, 28, 36, 1, 20, 22, 28, 36, 63, 70, 78, 84